

Molecular Biology
2.3 Carbohydrates and lipids

Diagrams showing the chemical structures of glucose and fructose, and the formation of sucrose from glucose and fructose.

Essential idea:

- Compounds of carbon, hydrogen and oxygen are used to supply and store energy.

Nature of Science:

- Evaluating claims
 - Health claims made about lipids in diets need to be assessed. (5.2)

The effects of "bad" LDL cholesterol compared to "good" HDL cholesterol

HW: Use **Evaluation Criteria** to evaluate an article on Saturated Fat, Trans-fat or Cholesterol

International-mindedness

- There is variation in the prevalence of different health problems around the world including obesity, dietary energy deficiency, kwashiorkor (protein malnutrition), anorexia nervosa and coronary heart disease.
- How do cultures affect the occurrences of these disorders.

Theory of knowledge:

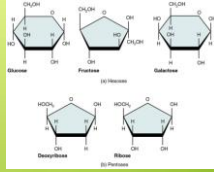
- There are conflicting views as to the harms and benefits of fats in diets.
- How do we decide between competing views?

Carbohydrates

- Monosaccharide monomers whose structures and bonding with each other by dehydration synthesis determine the properties and functions of the molecules

Monosaccharides

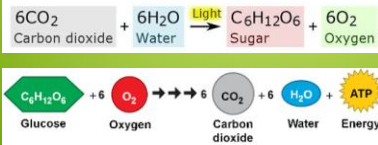
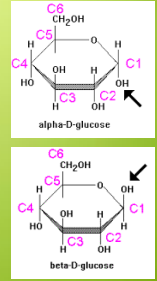
- CH₂O formula
- Simple single sugars
- Used for cellular respiration
- Aldoses and Ketoses



| | Triose sugars (C ₃ H ₆ O ₃) | Pentose sugars (C ₅ H ₁₀ O ₅) | Hexose sugars (C ₆ H ₁₂ O ₆) |
|----------------|---|--|--|
| Aldoses | <chem>C(C(CO)O)O</chem> Glyceraldehyde | <chem>C(C(C(CO)O)O)O</chem> Ribose | <chem>C(C(C(C(CO)O)O)O)O</chem> Glucose <chem>C(C(C(C(O)CO)O)O)O</chem> Galactose |
| Ketoses | | <chem>C(C(C(CO)O)O)C(=O)O</chem> Dihydroxyacetone <chem>C(C(C(CO)O)O)C(=O)O</chem> Ribulose | <chem>C(C(C(C(CO)O)O)C(=O)O)O</chem> Fructose |

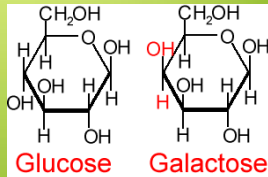
Monosaccharides

- Glucose (C₆H₁₂O₆)
 - Made by plants during photosynthesis
 - Metabolized during cellular respiration.
 - Main source of energy for plants and animals
 - Hexose ring structure in water (alpha and beta)
 - 5 of the carbons form corners on the ring with the 6th corner taken by oxygen



Monosaccharides

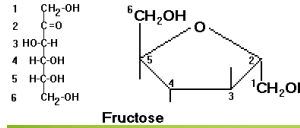
- Galactose
 - Found in mainly in milk
 - Hexose sugar
 - Less sweet as glucose
 - Converted by liver to glucose



What type of isomer of glucose is galactose?

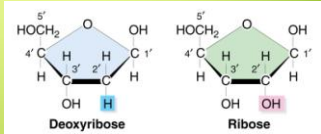
Monosaccharides

- Fructose
 - Pentose sugar
 - Found in fruit and honey.
 - Sweetest natural sugar.



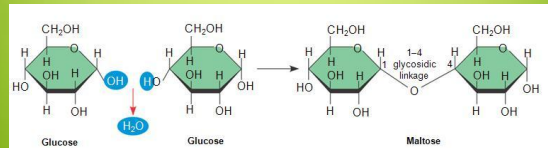
Monosaccharides

- Deoxyribose- DNA
- Ribose- RNA
- Both are pentose sugars
- Difference is at carbon 2






Disaccharides

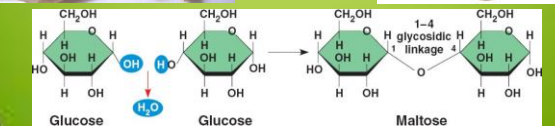
- Form by Glycosidic linkage
- Covalently bond by dehydration reaction



Disaccharides

- Maltose (C₁₂H₂₂O₁₂)
 - Glucose + Glucose dimer
 - Produced by hydrolysis of malt in grains








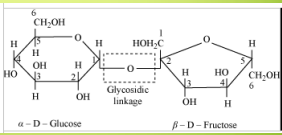


Glucose + Glucose → Maltose + H₂O

Disaccharides

- Sucrose (C₁₂H₂₂O₁₁)
 - Fructose + Glucose
 - Table sugar
 - From beets and cane

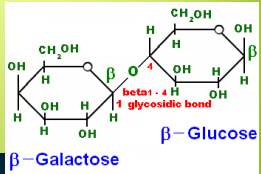







α-D-Glucose + β-D-Fructose → Sucrose

Disaccharides

- Lactose (C₁₂H₂₂O₁₁)
 - Glucose + Galactose
 - Found in milk

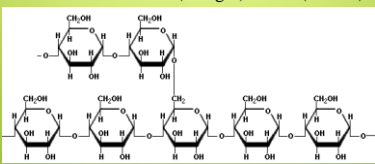


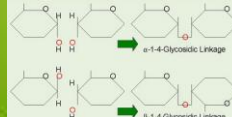
β-Galactose + α-Glucose → Lactose



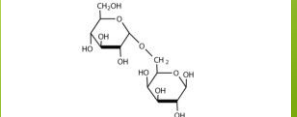
Polysaccharides

- Three or more monosaccharides
- Glycosidic bonds can be 1-4 (straight) or 1-6 (branch)






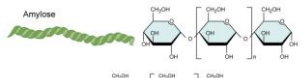
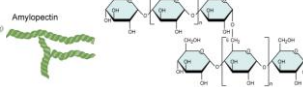
α-1,4-Glycosidic Linkage





β-1,4-Glycosidic Linkage

Polysaccharides


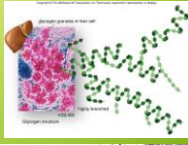
- Starch
 - The way plants store glucose (energy)
 - α-glucose linkages (spiral)
 - Amylose
 - un-branched and forms a helix
 - Amylopectin
 - chain is branched, so has a more globular shape

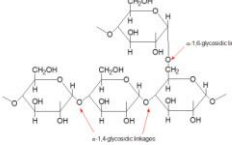




Polysaccharides

- Glycogen
 - 1-4 and 1-6 linkages on α-glucose
 - Creates spiral, branching and compaction
 - The way animals and some fungi store glucose.
 - In liver and muscle



Polysaccharides

- Cellulose
 - Cell walls of plants
 - 1-4 linkage of β -glucose molecules (linear)
 - Hydrogen bonds link the molecules together (microfibrils)
 - They have very high tensile strength.
 - Prevents plant cells from bursting, even under very high (water) pressure

Carbohydrate Directionality

- The nature of the bonding between carbohydrate subunits determines their relative orientation in the carbohydrate
- This determines the secondary structure of the carbohydrate.

| Polysacc | Monoacc | Bonds | Diagram |
|-----------------------|-------------------|---|---------|
| Starch: Amylose | α -glucose | 1-4 | |
| Starch: Amylopectin | α -glucose | 1-4 and 1-6 | |
| Glycogen (NOT starch) | α -glucose | 1-4 and 1-6 (more 1-6 than amylopectin) | |
| Cellulose | β -glucose | 1-4 | |

Lipids

- In general, lipids are nonpolar
- Don't dissolve in water
- More suitable for long term energy storage in humans than carbohydrates
- Fats, oils, waxes, phospholipids, steroids.

Fatty Acids

- The monomer that makes most lipids
 - hydrophilic carboxyl (COOH)
 - non-polar C-H bonds in fatty acid 'tails' (hydrophobic)
- Ester linkage: 3 fatty acids to 1 glycerol (condensation reaction)

Fatty Acid Structure

Fatty Acids

- Saturated (not Kinky)- only single bonds
- Unsaturated (Kinky)- some double bonds
 - Cis isomers
 - low melting points
 - usually liquid at room temperature
 - Trans isomers
 - high melting points
 - usually solid at room temperature

Cis- and Trans-Fatty Acids

Fatty Acids

- Polyunsaturated

| Properties | Cis Fatty acids | Trans Fatty acids |
|--|---|---|
| Health implications | Good | Detrimental; lowers good cholesterol and increases the level of bad cholesterol in the body. |
| Natural occurrence | Yes | Very less, hence produced artificially by partial hydrogenation of polyunsaturated fatty acids. |
| Orientation at double bonded carbon atoms | Hydrogen atoms at double bonded carbon atoms on same side | Hydrogen atoms at double bonded carbon atoms on opposite side |
| Recommended consumption | Can consume as per requirement | No more than 1% of total calories per day |
| Commonly found in | Natural fatty acids | Processed food, fast foods, butter and milky products. |
| Melting point | Lower | Higher |

Which are best?

